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12. The Shapes of
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Geometry Made Easy:

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Lewis Dot Structures

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*Valence shell electron
pair repulsion theory*

*13. Polar covalent
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11 Chap 4 | Chemical
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Answer key -

CHEMISTRY Lab

Report for VSEPR

Theory and Shapes of

Molecules Fill the

following tables. Do not

indicate polarity for

charged species (ions).

HCN 1. Lewis Structure

2. Perspective drawing

3. Number of atoms

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bonded to central atom
4. Number of non-
bonding electron pairs
on the central atom 5.

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The valence shell
electron pair repulsion
(VSEPR) theory (or
“VESPER” for short) is
how the geometry of a
molecule is determined
around a central atom.

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The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring the bond angles between the central atom and another atom bonded to it.

*Molecular Geometry
Vsepr Theory Worksheet
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VSEPR Theory. The VSEPR (Valence Shell Electron Pair Repulsion) model is used to predict the geometry of molecules based on the number of effective electron pairs around a central atom. The main postulate for the VSEPR theory is that the geometrical structure around a given atom is principally determined

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by minimizing the
repulsion between
effective electron pairs.

*17: VSEPR Theory and
Shapes of Molecules
(Experiment ...*

VSEPR Lab

Activity--ANSWER

KEY-2 - CHEM 1A

VSEPR Theory ...

Species Name: Lewis

Dot Structure:

Electronic Arrangement:

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BeF₂: linear: linear:
BCl₃: trigonal planar:
trigonal planar: CCl₄:
tetrahedral

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Molecules Fill the following tables. Do not indicate polarity for charged species (ions).

HCN 1. Lewis Structure

2. Perspective drawing

3. Number of atoms bonded to central atom

4. Number of non-bonding electron pairs on the central atom 5.

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LabAnswer key 4

0=6*6-3 §=C=:O. 2

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linear linear sp^3 $N \times 7 = -3$

μ a tetrahedral

Trpicpgoanmialdae sp^3

suis $B = 3$ Answer key -

CHEMISTRY The

valence shell electron

pair repulsion (VSEPR)

theory (or “VESPER”

for short) is how the

geometry of a molecule

is determined around a

central atom. The

molecular Page 4/25

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Explore molecule shapes by building molecules in 3D! How does molecule shape change with different numbers of bonds and electron pairs? Find out by adding single, double or triple bonds and lone pairs to the central atom. Then, compare the model to real molecules!

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*Molecule Shapes -
VSEPR / Lone Pairs /
Bonds - PhET ...*

Worksheet #1: Lewis
Structures Formula:
Lewis Structure:
Molecular Geometry
HBr linear

*VSEPR Worksheet 1
Answers*

The valence shell
electron pair repulsion

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(VSEPR) theory (or “VESPER” for short) is how the geometry of a molecule is determined around a central atom. The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring the bond angles between the central atom and another

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Molecular Shapes The

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shapes of molecules can be predicted from their Lewis structures by using the VSEPR (Valence Shell Electron Pair Repulsion) model, which states that electron pairs around a central atoms will ...

*Molecular Geometry
And Polarity Phet Lab
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VSEPR Theory: Shapes
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of Molecules -Part D.
When working on
VSEPR experiment: 1.
Completely answer all
questions and fill in all
blanks. 2. Draw all
Lewis structures. 3. If
present, show
nonbonding electron
pairs (or lone pairs) on
both central and non-
central atoms in Lewis
structures. 4.

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Chemistry 115 Lab - VSEPR Theory: Shapes of Molecules

VSEPR theory only predicts structure and cannot be used, by itself, to describe the places where electrons are allowed to be (i. e., the molecular orbitals). Valence Bond theory allows us to take a VSEPR structure (or a real structure) and get a

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rough idea of how the
electron density is
distributed in bond.

Molecular Modeling 1 / Chem Lab

Students will be able to
determine the shape of
molecules using VSEPR
theory as evidenced by
taking notes, performing
a molecule lab, and
doing whiteboards. Big
Idea Valence Shell

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Electron Pair Repulsion
Theory (VSEPR) allows
chemists to infer the
shape of molecules.

*Valence Shell Electron
Pair Repulsion Theory
(VSEPR)*

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Shell Electron Pair
Repulsion theory, or
VSEPR theory. The
following VSEPR table

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Vsepr Lab
supplies the names,
sketches and
descriptions of the most
common types of
molecular shapes that
you will encounter.
Note that several other
molecular geometries do
exist, however, they are
beyond the scope of this
course.

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Worksheet 13 -

Molecular Shapes The shapes of molecules can be predicted from their Lewis structures by using the VSEPR (Valence Shell Electron Pair Repulsion) model, which states that electron pairs around a central atoms will assume a geometry that keeps them as

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Worksheet 13 -

*Molecular Shapes Lewis
structures by using ...*

Teaching VSEPR model
theory? Utilize this

visual, active VSEPR

Molecular Geometry

Balloon introduction

lab, in print and digital

Google Apps format,

illustrating VSEPR 3D

molecular

shapes. Students begin

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this activity with a short
reading over electron
repulsion and draw
Lewis structures of 10
molecules and name
them. Balloons model
shapes illustrate shapes
and VSEPR shape
names.

*VSEPR Theory Model
Balloon Shapes Lab –
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VSEPR Theory: a
Page 33/37

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chemistry model used to predict the shape of individual molecules based on electron-pair electrostatic repulsion

VSEPR Model

The valence shell electron pair repulsion (VSEPR) model focuses on the bonding and nonbonding electron pairs present in the outermost (valence) shell of an atom that

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connects with two or
more other atoms.

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